

## CHEMISTRY

### CLASS-12

#### SOLUTIONS (*Osmotic pressure*)

**Q1.** If 1.71 g of sugar [*molar mass* = 342 ] are dissolved in 500 mL of an aqueous solution at 300K, what will be its osmotic pressure? **(Answer = 0.246 atm)**

**Q2.** A 10% aqueous solution of cane sugar [*mol. wt.* = 342 ] is isotonic with 1.754% aqueous solution of urea. Find the molecular mass of urea. **(Answer = 59.98)**

**Q3.** Define **osmosis** & **osmotic pressure**.